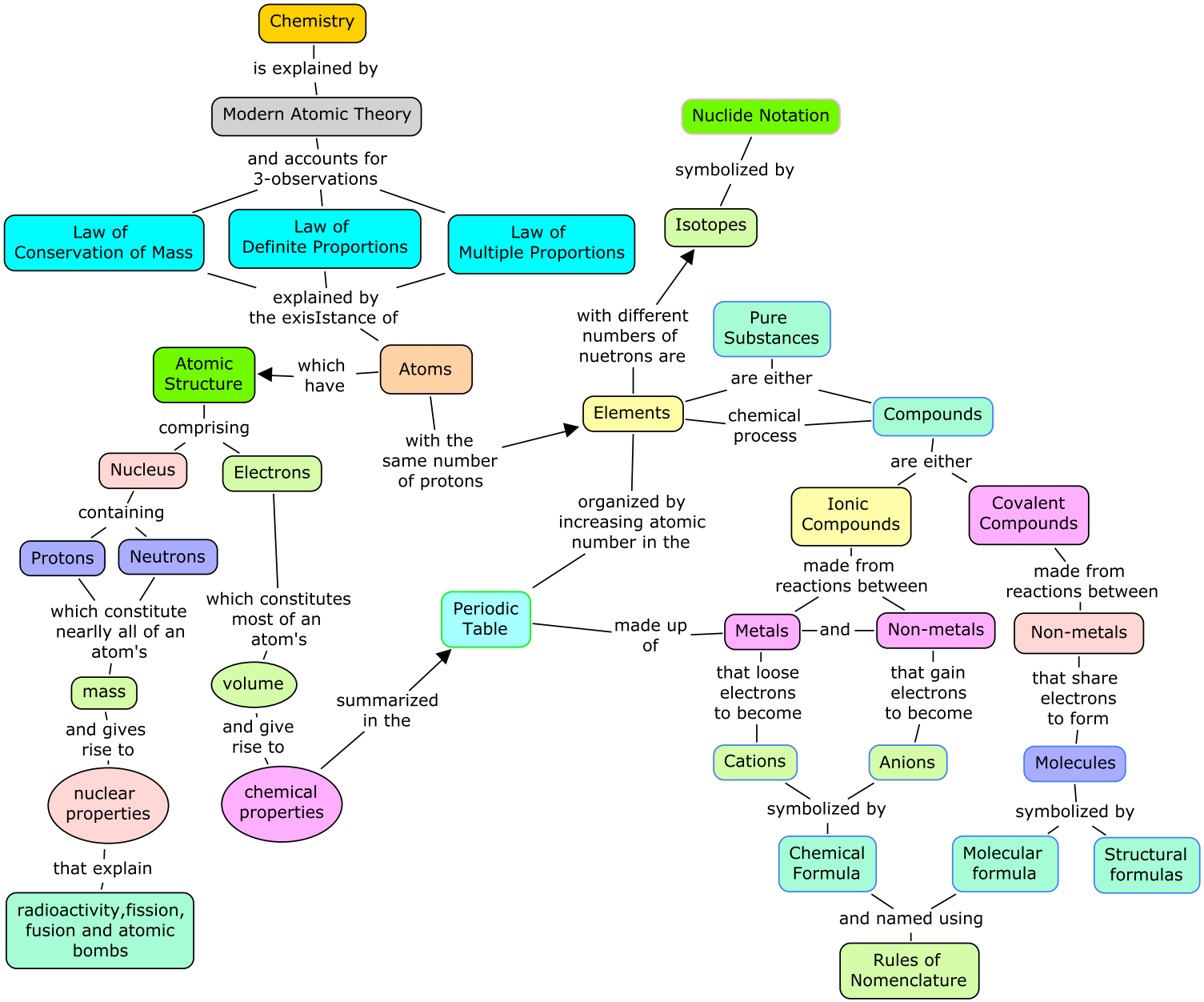
Chemistry
is explained by
Modern Atomic Theory
and accounts for 3-observations
Law of Conservation of Mass
Law of Definite Proportions
Law of Multiple Proportions
explained by the exisIstance of
Atoms
which have
Atomic Structure
comprising
Nucleus
Electrons
containing
Protons
Neutrons
which constitute nearlly all of an atom's
mass
and gives rise to
nuclear properties
that explain
radioactivity,fission, fusion and atomic bombs
which constitutes most of an atom's
volume
and give rise to
chemical properties
summarized in the
Periodic Table
with the same number of protons
Elements
with different numbers of nuetrons are
Isotopes
symbolized by
Nuclide Notation
organized by increasing atomic number in the
Pure Substances
are either
chemical process
Compounds
are either
Ionic Compounds
Covalent Compounds
made from reactions between
made from reactions between
made up of
Metals
and
Non-metals
Non-metals
that loose electrons to become
that gain electrons to become
that share electrons to form
Cations
Anions
Molecules
symbolized by
symbolized by
Chemical Formula
Molecular formula
Structural formulas
and named using
Rules of Nomenclature